Find ΔHrxn for the following **3 C(s) + 4 H2(g) 🡪 C3H8(g)**

Given: C3H8(g) + 5O2(g) 🡪 3CO2(g) + 4 H2O(g) ΔH = -2043 kJ

C(s) + O2(g) 🡪 CO2(g) ΔH = -393.5 kJ

2 H2(g) + O2(g) 🡪 2 H2O(g) ΔH = -483.6 kJ